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First Law of Thermodynamics. First Law of Thermodynamics: The heat energy supplied to the system increase the internal energy of the gas and the rest of the heat energy is converted into the work done by the system on the surroundings.  $dQ = dU + dW$ . where  $dQ$  is the given heat,  $dU$  is the change in internal energy, and  $dW$  is the work done by the system.

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A system which can exchange both energy and matter with its surroundings. A system which permits passage of energy but not mass, across its boundary. A system which can neither exchange energy nor matter with its surrounding. Part of the universe other than system, which can interact with it.

### JEE Chemistry Notes: Thermodynamics - Engineering

JEE : Thermodynamics, Chapter Notes, Class 11, Chemistry (IIT-JEE & AIPMT) JEE Notes | EduRev (1) Reversible isothermal expansion from A to B.  $DE AB = 0$  ,  $W AB = -nRT \ln (2)$  Reversible adiabatic expansion from B to C  $DE BC = nC V (T 2 -T 1)$   $W BC = DE BC (3)$  Isothermal compression from C to D  $DE CD ...$

### Thermodynamics, Chapter Notes, Class 11, Chemistry (IIT ...

This is a complete set of notes of Thermodynamics which is a part of Chemistry syllabus for NEET, JEE. Important notes of Chemistry for NEET, JEE for Thermodynamics are useful for all aspirants preparing for entrance exams including JEE, NEET. Important notes are also helpful for revision when you have less time and have to study many topics.

### Important Notes Of Chemistry For NEET, JEE: Thermodynamics ...

Revision Notes on Thermodynamics Thermodynamics:- It is the branch of physics which deals with process involving heat, work and internal energy. Thermodynamics is concerned with macroscopic behavior rather than microscopic behavior of the system.

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### JEE Mains: Thermodynamics - L 1 | IIT Chemistry | Unacademy JEE | Paaras Sir

SYNOPSIS. Thermodynamics is a branch of Chemistry which deals with the exchange in heat energy during chemical and physical changes. System: It is a part of the universe which is under thermodynamic consideration. Open system: Continuous exchange of heat energy and matter takes place together. Closed system: Continuous exchange of energy takes place, but matter remains constant.